Write your Centre number, candidate number and name on all the work you hand in.
Give details of the practical session and laboratory where appropriate, in the boxes provided.
Write in dark blue or black pen.
You may use an HB pencil for any diagrams or graphs.
Do not use staples, paper clips, glue or correction fluid.
DO NOT WRITE IN ANY BARCODES.

Answer all questions.
Electronic calculators may be used.
You may lose marks if you do not show your working or if you do not use appropriate units.
Use of a Data Booklet is unnecessary.

Qualitative Analysis Notes are printed on pages 10 and 11.
A copy of the Periodic Table is printed on page 12.

At the end of the examination, fasten all your work securely together.
The number of marks is given in brackets [ ] at the end of each question or part question.
1 Hydrogen peroxide, H₂O₂, can be oxidised to give oxygen, O₂. This reaction happens rapidly in the presence of acidified potassium manganate(VII), KMnO₄.

\[
2\text{MnO}_4^- (\text{aq}) + 5\text{H}_2\text{O}_2 (\text{aq}) + 6\text{H}^+ (\text{aq}) \rightarrow 2\text{Mn}^{2+} (\text{aq}) + 5\text{O}_2 (\text{g}) + 8\text{H}_2\text{O (l)}
\]

You will determine the concentration of a solution of hydrogen peroxide. You will first dilute the solution and then carry out a titration using acidified potassium manganate(VII), KMnO₄.

**FA 1** is aqueous hydrogen peroxide, H₂O₂.
**FA 2** is 0.0200 mol dm⁻³ potassium manganate(VII), KMnO₄.
**FA 3** is 1.0 mol dm⁻³ sulfuric acid, H₂SO₄.

(a) **Method**

**Dilution**
- Pipette 25.0 cm³ of **FA 1** into the 250 cm³ volumetric flask.
- Add distilled water to make 250 cm³ of solution and shake the flask thoroughly.
- Label this diluted solution of hydrogen peroxide **FA 4**.

**Titration**
- Fill the burette with **FA 2**.
- Rinse the pipette thoroughly with distilled water and then with a little **FA 4**.
- Pipette 25.0 cm³ of **FA 4** into a conical flask.
- Use the measuring cylinder to add 25 cm³ of **FA 3** to the conical flask.
- Perform a rough titration and record your burette readings in the space below.

The rough titre is ......................... cm³.

- Carry out as many accurate titrations as you think necessary to obtain consistent results.
- Make certain any recorded results show the precision of your practical work.
- Record in a suitable form below all of your burette readings and the volume of **FA 2** added in each accurate titration.

**Keep FA 3 and FA 4 for use in Question 2.**
(b) From your accurate titration results, obtain a suitable value for the volume of FA 2 to be used in your calculations. Show clearly how you obtained this value.

25.0 cm³ of FA 4 required ......................... cm³ of FA 2. [1]

(c) Calculations

Show your working and appropriate significant figures in the final answer to each step of your calculations.

(i) Calculate the number of moles of potassium manganate(VII) present in the volume of FA 2 calculated in (b).

moles of KMnO₄ = ......................... mol

(ii) Use your answer to (i) and the equation at the top of page 2 to calculate the number of moles of hydrogen peroxide present in 25.0 cm³ of FA 4.

moles of H₂O₂ = ......................... mol

(iii) Calculate the concentration, in mol dm⁻³, of H₂O₂ in FA 4.

concentration of H₂O₂ in FA 4 = ......................... mol dm⁻³

(iv) Calculate the concentration, in mol dm⁻³, of H₂O₂ in FA 1.

concentration of H₂O₂ in FA 1 = ......................... mol dm⁻³ [4]

[Total: 12]
In Question 1, hydrogen peroxide was oxidised to form oxygen. Hydrogen peroxide can also be reduced to form water. This occurs, for example, when hydrogen peroxide, H₂O₂, reacts with iodide ions, I⁻ (aq), to form iodine, I₂(aq).

\[ \text{H}_2\text{O}_2(\text{aq}) + 2\text{H}^+(\text{aq}) + 2\text{I}^-(\text{aq}) \rightarrow \text{I}_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l}) \]

The rate of this reaction can be measured by adding acidified hydrogen peroxide, H₂O₂, to a mixture of iodide ions, I⁻, thiosulfate ions, S₂O₃²⁻, and starch indicator. As the iodine is produced, it reacts immediately with the thiosulfate ions and is reduced back to iodide ions.

\[ \text{I}_2(\text{aq}) + 2\text{S}_2\text{O}_3^{2-}(\text{aq}) \rightarrow 2\text{I}^-(\text{aq}) + \text{S}_4\text{O}_6^{2-}(\text{aq}) \]

When all the thiosulfate has reacted, the iodine then turns the starch indicator blue-black. The rate of reaction may be determined by timing how long it takes the reaction mixture to turn blue-black.

In this experiment you will measure two reaction times and use one of these to calculate the concentration of the thiosulfate ions in FA 6.

FA 3 is 1.0 mol dm⁻³ sulfuric acid, H₂SO₄.
FA 4 is the diluted hydrogen peroxide, H₂O₂, you prepared in 1(a).
FA 5 is 1.0 mol dm⁻³ potassium iodide, KI.
FA 6 is aqueous sodium thiosulfate, Na₂S₂O₃.

starch indicator

Read through the instructions carefully before starting any practical work.

(a) Method

- Use the measuring cylinder to transfer 20 cm³ of FA 3 into the 100 cm³ beaker.
- Use the measuring cylinder to add 10 cm³ of FA 4 into the 100 cm³ beaker.
- Rinse the measuring cylinder with distilled water and drain.
- Use the measuring cylinder to transfer 20 cm³ of FA 5 into the 250 cm³ beaker.
- Use the measuring cylinder to transfer 20 cm³ of FA 6 into the 250 cm³ beaker.
- Use the measuring cylinder to transfer 10 cm³ of starch indicator into the 250 cm³ beaker.
- Add the contents of the 100 cm³ beaker to the 250 cm³ beaker and start timing immediately.
- Stir the mixture once and place the beaker on a white tile.
- Stop timing as soon as the solution turns blue-black.
- Record this reaction time to the nearest second.
- Rinse both beakers and drain.

Keep all FA solutions and starch indicator for use in (c).

reaction time = ......................... s  [2]
(b) Calculations

Show your working and appropriate significant figures in the final answer to each step of your calculations.

(i) The expression for the rate of reaction is shown.

\[ \text{rate} = \frac{\text{concentration of iodine at the reaction time}}{\text{reaction time}} \]

The concentration of iodine at the reaction time is the concentration of iodine that would have been present at the time the blue-black colour appeared if no thiosulfate ions had been added.

In this experiment, you should assume the rate = \(2.61 \times 10^{-5}\) mol dm\(^{-3}\) s\(^{-1}\).

Use this value for the rate to calculate the concentration of iodine that would have been present at the reaction time if no thiosulfate ions had been added.

\[ \text{concentration of } I_2 = \ldots \text{ mol dm}^{-3} \]

(ii) Use your answer to (i) and the total reaction volume, to calculate the number of moles of iodine that would have been present at the reaction time if no thiosulfate ions had been added.

(If you were unable to calculate an answer to (i), you should use the value of \(1.32 \times 10^{-3}\) mol dm\(^{-3}\). This may not be the correct value.)

\[ \text{moles of } I_2 = \ldots \text{ mol} \]

(iii) Calculate the number of moles of thiosulfate ions that reacted with the moles of \(I_2\), calculated in (ii).

\[ I_2(aq) + 2S_2O_3^{2-}(aq) \rightarrow 2I^- (aq) + S_4O_6^{2-}(aq) \]

\[ \text{moles of } S_2O_3^{2-} = \ldots \text{ mol} \]

(iv) Use your answer to (iii) to calculate the concentration of thiosulfate ions in FA 6.

\[ \text{concentration of } S_2O_3^{2-} \text{ in } FA 6 = \ldots \text{ mol dm}^{-3} \]

[4]
(c) Repeat the experiment in (a) using the following quantities of each reagent.

- add to the 100 cm³ beaker
  20 cm³ of FA 3
  10 cm³ of FA 4

- add to the 250 cm³ beaker
  20 cm³ of FA 5
  10 cm³ of FA 6
  10 cm³ of starch indicator
  10 cm³ of distilled water

- Add the contents of the 100 cm³ beaker to the 250 cm³ beaker and start timing immediately.
- Stir the mixture once and place the beaker on a white tile.
- Stop timing as soon as the solution turns blue-black.
- Record this reaction time to the nearest second.

reaction time = ............................. s  [1]

(d) (i) Explain the relationship between the value of the reaction time in (a) and the value of the reaction time in (c).

...............................................................................................................................................
...............................................................................................................................................
...............................................................................................................................................
...............................................................................................................................................

(ii) A student states that the error in the total volume of the reaction mixture in (a) is the same as the error in the total volume of the reaction mixture in (c).

State whether or not you agree with the student and explain your answer.

...............................................................................................................................................
...............................................................................................................................................
...............................................................................................................................................
...............................................................................................................................................
...............................................................................................................................................
...............................................................................................................................................

[4]

[Total: 11]
3 Qualitative Analysis

At each stage of any test you are to record details of the following.

- colour changes seen
- the formation of any precipitate
- the solubility of such precipitates in an excess of the reagent added

Where gases are released they should be identified by a test, described in the appropriate place in your observations.

You should indicate clearly at what stage in a test a change occurs. No additional tests for ions present should be attempted.

If any solution is warmed, a boiling tube MUST be used.

Rinse and reuse test-tubes and boiling tubes where possible.

Where reagents are selected for use in a test, the name or correct formula of the element or compound must be given.

(a) FA 7, FA 8, FA 9 and FA 10 are solutions that each contain a metal cation from those listed in the Qualitative Analysis Notes on page 10. Carry out tests using reagents that will allow you to identify which metal ions are present. Draw a single table to record your observations. Complete the table below with your conclusions.

<table>
<thead>
<tr>
<th>solution</th>
<th>FA 7</th>
<th>FA 8</th>
<th>FA 9</th>
<th>FA 10</th>
</tr>
</thead>
<tbody>
<tr>
<td>metal ion present</td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
</tbody>
</table>
(b) Carry out the following test and record your observations.

<table>
<thead>
<tr>
<th>test</th>
<th>observations</th>
</tr>
</thead>
<tbody>
<tr>
<td>To a 1 cm depth of FA 10 in a boiling tube add a 0.5 cm depth of 20 'vol' hydrogen peroxide, then add one drop of aqueous sodium hydroxide.</td>
<td></td>
</tr>
</tbody>
</table>

[2]

(c) FA 11 is a solid that contains one anion from those listed in the Qualitative Analysis Notes on page 11.

Place all of FA 11 into a boiling tube and add a 3 cm depth of distilled water. Stir until all the solid has dissolved.

(i) Carry out the following test and record your observations.

<table>
<thead>
<tr>
<th>test</th>
<th>observations</th>
</tr>
</thead>
<tbody>
<tr>
<td>To a 1 cm depth of the solution of FA 11 in a test-tube add dilute sulfuric acid.</td>
<td></td>
</tr>
</tbody>
</table>

(ii) Use your observations from (i) to identify the anion in FA 11.

The anion is ................... .

(iii) Use the remaining solution of FA 11 in the boiling tube to carry out a further test to support your identification of the anion. Record details of this test.

[4]

[Total: 17]
### Qualitative Analysis Notes

**Key:** \([ppt. = precipitate]\)

#### 1 Reactions of aqueous cations

<table>
<thead>
<tr>
<th>ion</th>
<th>reaction with NaOH(aq)</th>
<th>reaction with NH₃(aq)</th>
</tr>
</thead>
<tbody>
<tr>
<td>aluminium, (Al^{3+}(aq))</td>
<td>white ppt. soluble in excess</td>
<td>white ppt. insoluble in excess</td>
</tr>
<tr>
<td>ammonium, (NH₄⁺(aq))</td>
<td>no ppt. ammonia produced on heating</td>
<td>–</td>
</tr>
<tr>
<td>barium, (Ba^{2+}(aq))</td>
<td>no ppt. (if reagents are pure)</td>
<td>no ppt.</td>
</tr>
<tr>
<td>calcium, (Ca^{2+}(aq))</td>
<td>white ppt. with high ([Ca^{2+}(aq)])</td>
<td>no ppt.</td>
</tr>
<tr>
<td>chromium(III), (Cr^{3+}(aq))</td>
<td>grey-green ppt. soluble in excess</td>
<td>grey-green ppt. insoluble in excess</td>
</tr>
<tr>
<td>copper(II), (Cu^{2+}(aq))</td>
<td>pale blue ppt. insoluble in excess</td>
<td>blue ppt. soluble in excess giving dark blue solution</td>
</tr>
<tr>
<td>iron(II), (Fe^{2+}(aq))</td>
<td>green ppt. turning brown on contact with air insoluble in excess</td>
<td>green ppt. turning brown on contact with air insoluble in excess</td>
</tr>
<tr>
<td>iron(III), (Fe^{3+}(aq))</td>
<td>red-brown ppt. insoluble in excess</td>
<td>red-brown ppt. insoluble in excess</td>
</tr>
<tr>
<td>magnesium, (Mg^{2+}(aq))</td>
<td>white ppt. insoluble in excess</td>
<td>white ppt. insoluble in excess</td>
</tr>
<tr>
<td>manganese(II), (Mn^{2+}(aq))</td>
<td>off-white ppt. rapidly turning brown on contact with air insoluble in excess</td>
<td>off-white ppt. rapidly turning brown on contact with air insoluble in excess</td>
</tr>
<tr>
<td>zinc, (Zn^{2+}(aq))</td>
<td>white ppt. soluble in excess</td>
<td>white ppt. soluble in excess</td>
</tr>
</tbody>
</table>
2 Reactions of anions

<table>
<thead>
<tr>
<th>ion</th>
<th>reaction</th>
</tr>
</thead>
<tbody>
<tr>
<td>carbonate, CO$_3^{2-}$</td>
<td>CO$_2$ liberated by dilute acids</td>
</tr>
<tr>
<td>chloride, Cl$^-$ (aq)</td>
<td>gives white ppt. with Ag$^+$ (aq) (soluble in NH$_3$(aq))</td>
</tr>
<tr>
<td>bromide, Br$^-$ (aq)</td>
<td>gives cream ppt. with Ag$^+$ (aq) (partially soluble in NH$_3$(aq))</td>
</tr>
<tr>
<td>iodide, I$^-$ (aq)</td>
<td>gives yellow ppt. with Ag$^+$ (aq) (insoluble in NH$_3$(aq))</td>
</tr>
<tr>
<td>nitrate, NO$_3^-$ (aq)</td>
<td>NH$_3$ liberated on heating with OH$^-$ (aq) and Al foil</td>
</tr>
<tr>
<td>nitrite, NO$_2^-$ (aq)</td>
<td>NH$_3$ liberated on heating with OH$^-$ (aq) and Al foil; NO liberated by dilute acids (colourless NO $\rightarrow$ (pale) brown NO$_2$ in air)</td>
</tr>
<tr>
<td>sulfate, SO$_4^{2-}$ (aq)</td>
<td>gives white ppt. with Ba$^{2+}$ (aq) (insoluble in excess dilute strong acids)</td>
</tr>
<tr>
<td>sulfite, SO$_3^{2-}$ (aq)</td>
<td>gives white ppt. with Ba$^{2+}$ (aq) (soluble in excess dilute strong acids)</td>
</tr>
</tbody>
</table>

3 Tests for gases

<table>
<thead>
<tr>
<th>gas</th>
<th>test and test result</th>
</tr>
</thead>
<tbody>
<tr>
<td>ammonia, NH$_3$</td>
<td>turns damp red litmus paper blue</td>
</tr>
<tr>
<td>carbon dioxide, CO$_2$</td>
<td>gives a white ppt. with limewater (ppt. dissolves with excess CO$_2$)</td>
</tr>
<tr>
<td>chlorine, Cl$_2$</td>
<td>bleaches damp litmus paper</td>
</tr>
<tr>
<td>hydrogen, H$_2$</td>
<td>“pops” with a lighted splint</td>
</tr>
<tr>
<td>oxygen, O$_2$</td>
<td>relights a glowing splint</td>
</tr>
</tbody>
</table>
The Periodic Table of Elements

<table>
<thead>
<tr>
<th>Group</th>
<th>Periodic Table</th>
<th>Elements</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>1</td>
<td>H, He</td>
</tr>
<tr>
<td>2</td>
<td>2</td>
<td>Li, Be, B, C, N, O, F, Ne</td>
</tr>
<tr>
<td>3</td>
<td>3</td>
<td>Na, Mg, Al, Si, P, S, Cl, Ar</td>
</tr>
<tr>
<td>4</td>
<td>4</td>
<td>K, Ca, Sc, Ti, V, Cr, Mn, Fe, Co, Ni, Cu, Zn, Ga, Ge, As, Se, Br, Kr</td>
</tr>
<tr>
<td>5</td>
<td>5</td>
<td>Rb, Sr, Y, Zr, Nb, Mo, Tc, Ru, Rh, Pd, Ag, Cd, In, Sn, Sb, Te, I, Xe</td>
</tr>
<tr>
<td>6</td>
<td>6</td>
<td>Cs, Ba, La, Ce, Pr, Nd, Pm, Sm, Eu, Gd, Tb, Dy, Ho, Er, Tm, Yb, Lu</td>
</tr>
<tr>
<td>7</td>
<td>7</td>
<td>Fr, Ra, Ac, Th, Pa, U, Np, Pu, Am, Cm, Bk, Cf, Es, Fm, Md, No, Lr</td>
</tr>
</tbody>
</table>

**Key:***
- **atomic number**
- **atomic symbol**
- **name**
- **relative atomic mass**

**Group**
- **Periodic Table**
- **Elements**

**Elements Table**
- **Group 1:**
  - H: Hydrogen (1.0)
  - He: Helium (4.0)
- **Group 2:**
  - Li: Lithium (6.9)
  - Be: Beryllium (9.0)
- **Group 3:**
  - Na: Sodium (23.0)
  - Mg: Magnesium (24.3)
- **Group 4:**
  - K: Potassium (39.1)
  - Ca: Calcium (40.1)
- **Group 5:**
  - Rb: Rubidium (85.5)
  - Sr: Strontium (87.6)
- **Group 6:**
  - Cs: Cesium (122.9)
  - Ba: Barium (137.3)
- **Group 7:**
  - Fr: Francium (288.0)
  - Ra: Radon (226.0)
- **Group 8:**
  - Ac: Actinium (227.0)
  - Th: Thorium (232.0)

**Notes:**
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